**Decomposition reactions (Level 1) examiners tips: Read these please!**

**•** 1 largish compound 🡪 2 or 3 smaller compounds, one of which is usually a gas

**•** decomposition using heat aka thermal decomposition

*eg metal carbonate 🡪 metal oxide + carbon dioxide*

**•** decomposition which occurs naturally but is sped up using a catalyst aka catalytic decomposition

*eg hydrogen peroxide using a catalyst (manganese dioxide or potassium iodide or potato or liver)*

**•** positive test for water: cobalt chloride paper turns from a blue to a pink colour

**•** positive test for carbon dioxide: bubble the gas into limewater which turns milky

or

CO2 gas will extinguish a flame

**Also…”don’t be daft”**

the symbol of oxygen is a CAPITAL O, this also applies for in nitrates NO3-, carbonates CO32- sulfates SO42-

be careful, Na2CO3 (sodium carbonate) does NOT decompose thermally

a catalyst is a chemical that SPEEDS UP the rate of a reaction

a catalyst is NOT USED UP during the reaction.

seeing BUBBLES is an observation, a gas forms is not.

how would you see a colourless gas eg water vapour, usually by observing CONDENSATION on a tube

an observation of “seeing steam” or “a colourless liquid” is NOT acceptable

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